

KIPSONLINEEARLYPREPSESSION

BY SAEED MDCAT

www.saeedmdcat.com
FB/SAEED MDCAT
03418729745



SAEED MDCAT
SAEED MDCAT TEAM
f SAEEDMDCAT



SAEED MDCAT
SAEED MDCAT TEAM
f SAEEDMDCAT

SAEED MDCAT TEAM

f SAEEDMDCAT

SAEED MDCAT

TRUE HERO ALWAYS WINS

National MDCAT 2020 Results

190+ 35 Students

185+ 218 Students

180+ 677 Students



SAEED MDCAT
SAEED MDCAT TEAM
f SAEEDMDCAT

SAEED MDCAT
JOIN US FOR FREE
TO SECURE YOUR FUTURE
SAEED MDCAT TEAM

Designed by: Usman Arts

f SAEEDMDCAT



SAEED MDCAT

SAEED MDCAT TEAM

f SAEEDMDCAT

f SAEEDMDCAT

TRUE HERO ALWAYS WINS**SAEED MDCAT**
National MDCAT 2020 Results**190+****35 Students****185+****218 Students****180+****677 Students****SAEED MDCAT**

Education



Like

**Follow**

8,592 people like this

Home

Reviews

About

Videos

Photos

Posts

About

Suggest edits

<http://www.saeedmdcat.com/>

Send message



0341 8729745



+92 347 1729745



Education

See All >

SAEED MDCAT

If you want to get all the test and Lec of all
Academy then follow us on different
platforms

WEBSITE.www.saeedmdcat.com

FB Group And Page.SAEED MDCAT

INSTAGRAM.SAEED MDCAT

Twitter.Smdcat

For WhatsApp Group.03418729745

Regards.Huzaiifa Saeed,Usama Sohail

Fizza Maryum,Javeria Shukor

SAEED MDCAT TEAM



SAEEDMDCAT

www.saeedmdcat.com



FB) SAEED MDCAT

03418729745



QUIZZES

Practice Test-1(Introduction
to Fundamental Concepts of Chem...



10 Questions



7 min

Topics

Atomic mass

Start Quiz

SAEED MDCAT

SAEED MDCAT TEAM



SAEEDMDCAT

06 : 36



2/10



7 min



Hint

Q : Which of the following term is used for the mass of chlorine 35.5 amu

A

Relative atomic mass

B

Relative Molecular mass

C

Mass number

D

Relative isotopic mass

SAEED MDCAT

SAEED MDCAT TEAM



SAEEDMDCAT

www.saeedmdcat.com

FB)SAEED MDCAT

03418729745

1

2

3

4

5

6

7

06 : 34



3/10



7 min



Hint

Q : Which of the following statement is wrong about isotopes



They possess different mass number



They possess same chemical properties



They possess different physical properties



They possess different position in the periodic table

SAEED MDCAT

SAEED MDCAT TEAM



SAEEDMDCAT

www.saeedmdcat.com

FB)SAEED MDCAT

03418729745

1

2

3

4

5

6

7

06 : 31



4/10



7 min



Hint

Q : Boron has two stable isotopes, ^{10}B (19%) and ^{11}B (81%) Find the average atomic mass of Boron

A

$$\frac{(19 \times 10) + (81 \times 10)}{100}$$

B

$$\frac{(19 \times 10) + (81 \times 11)}{100}$$

C

$$\frac{(81 \times 10) + (19 \times 11)}{100}$$

D

$$\frac{(19 \times 10) \times (81 \times 11)}{100}$$

SAEED MDCAT

SAEED MDCAT TEAM



SAEEDMDCAT

www.saeedmdcat.com

FB) SAEED MDCAT

1

2

3

4

5

6

7

03418729745

06 : 29



5/10



7 min



Hint

Q : In nature the ratio of relative percentage abundance of the isotopes ^{35}Cl and ^{37}Cl is _____ respectively



1 : 3



3 : 1



1 : 1



1 : 4

SAEED MDCAT

SAEED MDCAT TEAM



SAEEDMDCAT

www.saeedmdcat.com

FB)SAEED MDCAT

03418729745

1

2

3

4

5

6

7

Q : 1 amu is equal to

- ☐ $1.661 \times 10^{-27} \text{kg}$
- ☐ $1.661 \times 10^{-24} \text{g}$
- ☐ $1.661 \times 10^{-21} \text{mg}$
- ☐ All of these

SAEED MDCAT

SAEED MDCAT TEAM

SAEEDMDCAT

www.saeedmdcat.com

FB)SAEED MDCAT

03418729745

Q : The atomic weight of B is 10.8. There are only two naturally occurring isotopes of boron and . The natural abundance of the isotope must be

☐ 10%

☐ 20%

☐ 50%

☐ 80%

SAEED MDCAT

SAEED MDCAT TEAM

SAEEDMDCAT

www.saeedmdcat.com

FB)SAEED MDCAT

03418729745

Q : Total number of radioactive isotopes produced through artificial disintegration

A 240

B 40

C 300

D 340

SAEED MDCAT

SAEED MDCAT TEAM

SAEEDMDCAT

www.saeedmdcat.com

FB)SAEED MDCAT

03418729745

4

5

6

7

9

10

Q : Which pair of elements have same number of isotopes

- ☐ A B, Cl
- ☐ B C, I
- ☐ C Na, Ca
- ☐ D Ni, Cd

SAEED MDCAT

SAEED MDCAT TEAM

SAEEDMDCAT

www.saeedmdcat.com

FB)SAEED MDCAT

03418729745

Q :

The atomic weight of Cu is 63.546. There are only two naturally occurring isotopes of copper, ^{63}Cu and ^{65}Cu . The natural abundance of the ^{63}Cu isotope must be

10%

30%

50%

70%

SAEED MDCAT

SAEED MDCAT TEAM

SAEEDMDCAT

www.saeedmdcat.com

FB)SAEED MDCAT

03418729745

WWW.SAEEDMDCAT.COM

FB) SAEED MDCAT

03415729745

QUIZZES

Practice Test-2 (Introduction
to Fundamental Concepts of Chem...



2 min



1 min

001

11:11:51

SAEED MDCAT

SAEED MDCAT TEAM



SAEEDMDCAT

Q : One mole of Hydrogen and oxygen have same at STP

- ☐ A Gram molecular weight
- ☐ B Protons in the molecules
- ☐ C molar volume
- ☐ D Electrons in the valence shell

SAEED MDCAT

SAEED MDCAT TEAM

SAEEDMDCAT



2/16

Time



Unit

Q : 18.0g of glucose contains number of hydrogen atoms



7.2×10^{23}



6.3×10^{23}



2.7×10^{23}



3.6×10^{23}

SAEED MDCAT

SAEED MDCAT TEAM



SAEEDMDCAT



Q : One mole of H_2SO_4 contain

- ☐ A 6.022×10^{23} number of hydrogen atoms
- ☐ B 3.011×10^{23} atoms of Sulphur
- ☐ C 2.4088×10^{24} atoms of oxygen
- ☐ D 2.4088×10^{24} molecules of oxygen

SAEED MDCAT

SAEED MDCAT TEAM

SAEEDMDCAT



4/10

Time



Exit

Q : 4g H_2 reacts with 32.0g O_2 to produce water. Which of the following statements is correct



H_2 -limiting reactant



O_2 -non-limiting reactant



2.0 mole water is produced



1 mole water is produced

SAEED MDCAT

SAEED MDCAT TEAM



SAEEDMDCAT





Q (10)

Time



Ans

Q : 2.8g of N_2 molecule contains number of chemical bonds



6.02×10^{22}



1.204×10^{23}



1.8×10^{23}



1.8×10^{22}

SAEED MDCAT

SAEED MDCAT TEAM



SAEEDMDCAT



Q : The number of moles of KMnO_4 that contain 1 mole of oxygen

- ☐ 2 moles
- ☐ 0.5 moles
- ☒ 0.25 moles
- ☐ 1.5 moles

SAEED MDCAT

SAEED MDCAT TEAM

SAEEDMDCAT

Q : The number of atoms in one gram atom of an element is

- ☐ A N_A of atoms
- ☐ B N_A of ions
- ☐ C N_A of molecules
- ☐ D N_A of formula unit

SAEED MDCAT

SAEED MDCAT TEAM

SAEEDMDCAT

Q : One gram molecule of different gases have all the following properties same at STP except

- ☐ Molecules
- ☐ Moles
- ☐ Volume
- ☐ Masses

SAEED MDCAT

SAEED MDCAT TEAM

SAEEDMDCAT

Q : The total number of O-atoms in 18g of glucose are

- ☐ A 6.02×10^{22}
- ☐ B 6.02×10^{23}
- ☐ C 3.6×10^{23}
- ☐ D 3.6×10^{22}

SAEED MDCAT

SAEED MDCAT TEAM

SAEEDMDCAT

Q : How many moles of neutron are present in one mole of heavy water

☐ 10

☐ 18

☐ 8

☐ 20

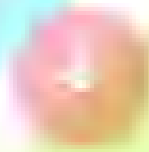
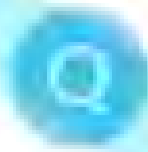
SAEED MDCAT

SAEED MDCAT TEAM

SAEEDMDCAT

QUIZZES

Practice Test-3(Introduction
to Fundamental Concepts of Chem...



SAEED MDCAT

SAEED MDCAT TEAM



SAEEDMDCAT

www.saeedmdcat.com

FB)SAEED MDCAT

03418729745

Q : The volume occupied by 1.6g of O_2 at STP is

☐ 2.24dm³

☐ 22.4dm³

☐ 1.12dm³

☐ 112cm³

SAEED MDCAT

SAEED MDCAT TEAM

SAEEDMDCAT

www.saeedmdcat.com

FB)SAEED MDCAT

03418729745



2/16



1 min



100%

Q : Eight grams of methane occupies volume at STP



22.4 dm³



2.24 dm³



1.12 dm³



11.2 dm³

SAEED MDCAT

SAEED MDCAT TEAM



SAEEDMDCAT

www.saeedmdcat.com

FB)SAEED MDCAT

03418729745

Q : Avogadro's number is the number of molecules present in

- ☐ A 1 dm^3 of molecule
- ☐ B 1 g of formula mass
- ☐ C Gram molecular mass
- ☐ D 1 g of atom

SAEED MDCAT

SAEED MDCAT TEAM

SAEEDMDCAT

www.saeedmdcat.com

FB)SAEED MDCAT

03418729745

Q:

15 gram of a gas occupies 11.2 dm^3 at S.T.P,
the gas is

☐ CO

☐ NO

☐ CO_2

☐ N_2O

SAEED MDCAT

SAEED MDCAT TEAM

SAEEDMDCAT

www.saeedmdcat.com

FB)SAEED MDCAT

03418729745

Q : 4 g of CH_4 gas has molar volume at S.T.P

☐ A 22.414 cm^3

☐ B 5.60 dm^3

☐ C 11.2 cm^3

☐ D 22414 cm^3

SAEED MDCAT

SAEED MDCAT TEAM

SAEEDMDCAT

www.saeedmdcat.com

FB)SAEED MDCAT

03418729745

Q : Hydrogen and oxygen have same at STP

- ☐ Gram molecular weight
- ☐ Protons in the molecules
- ☐ Gram molecular volume
- ☐ Electrons in the valence shell

SAEED MDCAT

SAEED MDCAT TEAM

SAEEDMDCAT

www.saeedmdcat.com

FB)SAEED MDCAT

03418729745

Q : Equal volumes of N_2O and CO_2 are taken in identical conditions, the correct relation between the masses of two gases is

☐ $\text{N}_2\text{O} > \text{CO}_2$

☒ $\text{N}_2\text{O} < \text{CO}_2$

☐ $\text{N}_2\text{O} = \text{CO}_2$

☐ $\text{N}_2\text{O}^3 \text{CO}_2$

SAEED MDCAT

SAEED MDCAT TEAM

SAEEDMDCAT

www.saeedmdcat.com

FB)SAEED MDCAT

03418729745

Q : Which one is incorrect relation at STP

- ☐ A 6g of carbon = 3.01×10^{23} atoms
- ☐ B 11.2 dm³ of CO₂ = 3.01×10^{23} molecules
- ☐ C 49 g of H₂SO₄ = 4 moles of atoms
- ☐ D 1 mole of sucrose = 45 moles of atoms

SAEED MDCAT

SAEED MDCAT TEAM

SAEEDMDCAT

www.saeedmdcat.com

FB)SAEED MDCAT

03418729745

Q : Avogadro's number is the number of molecules present in

- ☐ A 1 dm^3 of molecule
- ☐ B 1 g of hydrogen gas
- ☐ C 1 g atom
- ☐ D Gram molecular mass

SAEED MDCAT

SAEED MDCAT TEAM

SAEEDMDCAT

www.saeedmdcat.com

FB)SAEED MDCAT

03418729745

Q : The volume of oxygen gas is 1.12dm^3 at STP , the mass of oxygen approximately is

☐ 3.2g

☐ 1.6g

☐ 2.8g

☐ 1.4g

SAEED MDCAT

SAEED MDCAT TEAM

SAEEDMDCAT

www.saeedmdcat.com

FB)SAEED MDCAT

03418729745

www.saeedmdcat.com

FB) SAEED MDCAT

03418729745

QUIZZES

Practice Test-4 (Introduction
to Fundamental Concepts of Chem...



20 min



7 min

100%

TEST END

SAEED MDCAT

SAEED MDCAT TEAM



SAEEDMDCAT

Q : Which of the following term is not used for ionic compounds?

- ☐ A Formula unit
- ☐ B Empirical formula
- ☐ C Molecular formula
- ☐ D Formula Mass

SAEED MDCAT

SAEED MDCAT TEAM

SAEEDMDCAT

Q : Which one of the following statements is not involved in the determination of empirical formula

- ☐ A %age of each element
- ☐ B gram atom of each element
- ☐ C isotopes of each element
- ☐ D atomic ratio of element

SAEED MDCAT

SAEED MDCAT TEAM

SAEEDMDCAT



Q : Which is correct statement

- ☐ Rectified spirit is 100% ethanol
- ☐ Molar volume of an ideal gas is 24 dm^3 at STP
- ☐ R_f value for red ink has units of cm^{-1}
- ☐ Quantitative analysis involves four steps

SAEED MDCAT

SAEED MDCAT TEAM

SAEEDMDCAT

Q : Elemental analysis is performed to determine

- ☐ Molar mass of the compound
- ☐ Structural formula of a compound
- ☐ Empirical formula of a compound
- ☐ Mass of halogen present in a compound

SAEED MDCAT

SAEED MDCAT TEAM

SAEEDMDCAT

Q :

There are different steps in determining the empirical formula

Step I. Calculating the number of gram atom

Step II. Determining the atomic ratio

Step III. Determining the percentage composition

What is the correct sequence of the above steps?



I, II, III



III, II, I



II, I, III



III, I, II

Q : The simplest formula of a compound containing 50% of element X (At.wt = 10) and 50% of element Y (At. wt = 20) is

- ☐ A XY
- ☐ B XY_2
- ☐ C X_2Y
- ☐ D X_2Y_3

SAEED MDCAT

SAEED MDCAT TEAM

SAEEDMDCAT

Q : While determine molecular formula, the simple multiple 'n' is not unity for

- ☐ A H_2O
- ☐ B H_2O_2
- ☐ C $C_{12}H_{22}O_{11}$
- ☐ D CO_2

SAEED MDCAT

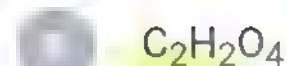
SAEED MDCAT TEAM

SAEEDMDCAT

FB) SAEED MDCAT

03418729745

Q : An acid with molecular mass 104 contains 34.6% C, 3.85% H and rest is O. The molecular formula of acid is



SAEED MDCAT

SAEED MDCAT TEAM

SAEEDMDCAT

FB) SAEED MDCAT

03418729745

Q : A compound contains 50% S and 50% O by mass. The empirical formula of compound is

☐ SO_2

☐ S_2O_3

☐ SO_3

☒ SO

SAEED MDCAT

SAEED MDCAT TEAM

SAEEDMDCAT

Q : A pair of compounds that has same empirical formula

- ☐ A Acetic acid and glucose
- ☐ B Acetic acid and formic acid
- ☐ C Formic acid and sucrose
- ☐ D Ethane and Ethyne

SAEED MDCAT

SAEED MDCAT TEAM

SAEEDMDCAT

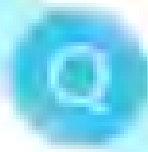
www.saeedmdcat.com

FB) SAEED MDCAT

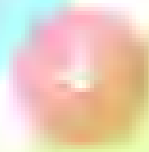
03418729745

QUIZZES

Practice Test-5 (Introduction
to Fundamental Concepts of Chem...



2 min 21 s



7 min

001

11:11 PM

SAEED MDCAT

SAEED MDCAT TEAM



SAEEDMDCAT

Q : For a 10% solution of acetic acid in water contain 1.850 mole is the calculate value of

- ☐ Normality
- ☐ Molarity
- ☒ Molality
- ☐ Formality

SAEED MDCAT

SAEED MDCAT TEAM

SAEEDMDCAT

www.saeedmdcat.com

FB)SAEED MDCAT

03418729745

Q : A solution of glucose is 10% W/V. The volume in which 1g mole of it is dissolved will be

- ☐ A 1dm^3
- ☐ B 1.8dm^3
- ☐ C 200cm^3
- ☐ D 900cm^3

SAEED MDCAT

SAEED MDCAT TEAM



SAEEDMDCAT

www.saeedmdcat.com

FB)SAEED MDCAT

03418729745

Q : A solution of 92g of ethanol, 96g methanol and 90g water has mole fraction of ethanol equal to

☐ 0.1

☐ 0.4

☐ 0.2

☐ 0.5

SAEED MDCAT

SAEED MDCAT TEAM



SAEEDMDCAT

www.saeedmdcat.com

FB)SAEED MDCAT

03418729745

Q : Which of the following solution pairs can be separated into its pure components by fractional distillation?

☐ A Benzene – toluene

☐ B Water -HNO₃

☐ C Water-HCl

☐ D Water-C₂H₅OH

SAEED MDCAT

SAEED MDCAT TEAM

SAEEDMDCAT

www.saeedmdcat.com

FB)SAEED MDCAT

03418729745

Q : If 5.85 of NaCl are dissolved in 90g of water, the mole fraction of solute is

- ☐ 0.01
- ☐ 0.025
- ☐ 0.0196
- ☐ 0.0382

SAEED MDCAT

SAEED MDCAT TEAM

SAEEDMDCAT

www.saeedmdcat.com

FB)SAEED MDCAT

03418729745

Q : If 5.85 g of NaCl are dissolved in 90g of water, the mole fraction of NaCl is:

- ☐ 0.1
- ☐ 0.01
- ☐ 0.2
- ☐ 0.0196

SAEED MDCAT

SAEED MDCAT TEAM

SAEEDMDCAT

www.saeedmdcat.com

FB)SAEED MDCAT

03418729745

Q : The amount of NaOH required to make 1dm^3 of 0.5M aqueous solution

- ☐ 2g
- ☐ 10g
- ☐ 4g
- ☐ 20g

SAEED MDCAT

SAEED MDCAT TEAM

SAEEDMDCAT

www.saeedmdcat.com

FB)SAEED MDCAT

03418729745

Q : Number of moles of solute present in 50cm^3 of 0.1M aqueous solution is

- ☐ 0.05
- ☐ 0.005
- ☐ 0.001
- ☐ 0.1

SAEED MDCAT

SAEED MDCAT TEAM

SAEEDMDCAT

www.saeedmdcat.com

FB)SAEED MDCAT

03418729745

4

5

6

7

9

10

Q : An aqueous solution of 0.05M NaOH is available. What is %age W/V of this solution.

- ☐ A 2%
- ☐ B 0.2%
- ☐ C 4%
- ☐ D 10%

SAEED MDCAT

SAEED MDCAT TEAM

SAEEDMDCAT

www.saeedmdcat.com

FB)SAEED MDCAT

03418729745

Q : A solution of glucose is 10% W/V. The volume in which 1g mole of it is dissolved will be

- ☐ A 1dm^3
- ☐ B 1.8dm^3
- ☐ C 200cm^3
- ☐ D 900cm^3

SAEED MDCAT

SAEED MDCAT TEAM



SAEEDMDCAT

www.saeedmdcat.com

FB)SAEED MDCAT

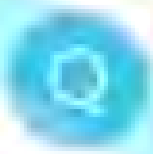
03418729745

 SAEEDMDCAT

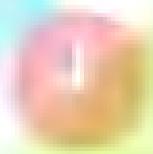


QUIZZES

Test Unit-1(Introduction to
Fundamental Concepts of Chemist...



010 — 000



00 — 000

000

SAEED MDCAT

SAEED MDCAT TEAM



SAEEDMDCAT

www.saeedmdcat.com

FB)SAEED MDCAT

03418729745

Q : Haemoglobin molecule is _____ times
then heavier H_2

10,000

68,000

1000

34,000

SAEED MDCAT

SAEED MDCAT TEAM

SAEEDMDCAT

www.saeedmdcat.com

FB)SAEED MDCAT

03418729745

Q:

The isotopes $^{16}_8\text{O}$, $^{24}_{12}\text{Mg}$, $^{28}_{14}\text{Si}$, $^{40}_{20}\text{Ca}$, $^{56}_{26}\text{Fe}$ form nearly _____ of earth crust

☐ 20%

☐ 30%

☐ 40%

☐ 50%

SAEED MDCAT

SAEED MDCAT TEAM

SAEEDMDCAT

www.saeedmdcat.com

FB)SAEED MDCAT

03418729745



Q (10)



20 min



Exit

Q : Which one of the following pair is isoelectronic



Si, CO_2



H_2O , Ne



Na, K



NH_3 , Mg

SAEED MDCAT

SAEED MDCAT TEAM



SAEEDMDCAT

www.saeedmdcat.com

FB) SAEED MDCAT

03418729745

Q : Which one has least number of isotopes

- ☐ Tin
- ☐ Cadmium
- ☐ Hydrogen
- ☐ Palladium

SAEED MDCAT

SAEED MDCAT TEAM

SAEEDMDCAT

www.saeedmdcat.com

FB)SAEED MDCAT

03418729745

Q : The number of moles of CH_4 which contains 3.0g of Carbon

- ☐ 1.0
- ☐ 0.75
- ☐ 0.5
- ☐ 0.25

SAEED MDCAT

SAEED MDCAT TEAM

SAEEDMDCAT

www.saeedmdcat.com

FB)SAEED MDCAT

03418729745

Q : A pair of compounds that has same empirical formula

- ☐ A Acetic acid and glucose
- ☐ B Acetic acid and formic acid
- ☐ C Formic acid and sucrose
- ☐ D Both a and b

SAEED MDCAT

SAEED MDCAT TEAM

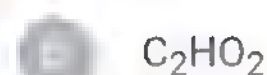
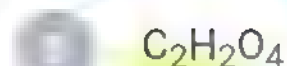
SAEEDMDCAT

www.saeedmdcat.com

FB)SAEED MDCAT

03418729745

Q : An acid with molecular mass 104 contain 34.6%C, 3.85% H and rest is O the molecular formula of acid is



SAEED MDCAT

SAEED MDCAT TEAM

SAEEDMDCAT

www.saeedmdcat.com

FB)SAEED MDCAT

03418729745

Q : CH_2O is the empirical formula of

- ☐ $\text{C}_{12}\text{H}_{22}\text{O}_{11}$
- ☐ $\text{CH}_2(\text{OH})\text{CH}_2(\text{OH})$
- ☐ $\text{CH}_3 \cdot \text{CH}(\text{OH}) \cdot \text{COOH}$
- ☐ CH_3CHO

SAEED MDCAT

SAEED MDCAT TEAM

SAEEDMDCAT

www.saeedmdcat.com

FB)SAEED MDCAT

03418729745

6

7

8

9

10

11

12

Q : The simplest formula of a compound containing 50% of element X (At.wt = 10) and 50% of element Y (At. wt = 20) is

- ☐ A XY
- ☐ B XY_2
- ☐ C X_2Y
- ☐ D X_2Y_3

SAEED MDCAT

SAEED MDCAT TEAM

SAEEDMDCAT

www.saeedmdcat.com

FB)SAEED MDCAT

03418729745

intan

20 min

100%

Q : 3.0 g of NO gas occupies volume

☐ 22.424 dm³

☐ 2.2414 dm³

☐ 11.2 dm³

☐ 1.2 dm³

SAEED MDCAT

SAEED MDCAT TEAM

SAEEDMDCAT

www.saeedmdcat.com

FB)SAEED MDCAT

03418729745

6

7

8

9

11

12

Q : One gram molecule of different gases have all the following properties same at STP except

- ☐ A Molecules
- ☐ B Moles
- ☐ C Volume
- ☐ D Masses

SAEED MDCAT

SAEED MDCAT TEAM

SAEEDMDCAT

www.saeedmdcat.com

FB)SAEED MDCAT

03418729745

Q : 720g of water contains how many moles of water?

☐ A 20

☐ B 60

☐ C 40

☐ D 80

SAEED MDCAT

SAEED MDCAT TEAM

SAEEDMDCAT

www.saeedmdcat.com

FB)SAEED MDCAT

03418729745

Q : An atom of carbon is twelve times heavier than ____ atom

- ☐ H
- ☐ Ne
- ☐ He
- ☐ Li

SAEED MDCAT

SAEED MDCAT TEAM

SAEEDMDCAT

www.saeedmdcat.com

FB)SAEED MDCAT

03418729745

MCQ 20 20 min 100%

Q : What is the volume (in dm^3) of CO_2 liberated at STP, when 53gram of sodium carbonate (mol. mass = 106) is treated with excess dilute HCl in following reaction

- ☐ 11.2
- ☐ 0.448
- ☐ 22.414
- ☐ 5.51

SAEED MDCAT

SAEED MDCAT TEAM

SAEEDMDCAT

www.saeedmdcat.com

FB)SAEED MDCAT

03418729745

Q : 500 ml of NH_3 contains 6.00×10^{23} molecules at S.T.P. How many molecules are present in 100 ml of CO_2 at S.T.P?

☐ 6×10^{23}

☒ 1.5×10^{23}

☐ 1.2×10^{23}

☐ 3.01×10^{23}

SAEED MDCAT

SAEED MDCAT TEAM

SAEEDMDCAT

www.saeedmdcat.com

FB)SAEED MDCAT

03418729745

Q : Avogadro's number is the number of molecules present in

- ☐ A 11.2 dm³ at STP
- ☐ B 22400 cm³ at STP
- ☐ C 1000 cm³ at STP
- ☐ D 1 cm³ at STP

SAEED MDCAT

SAEED MDCAT TEAM

SAEEDMDCAT

www.saeedmdcat.com

FB)SAEED MDCAT

03418729745

Q : What volume is occupied by a mixture of 0.5g H_2 , 16g O_2 and 7.0g N_2

☐ 2.24 dm^3

☐ 22.4 dm^3

☐ 0.224 dm^3

☐ 11.2 dm^3

SAEED MDCAT

SAEED MDCAT TEAM

SAEEDMDCAT

www.saeedmdcat.com

FB)SAEED MDCAT

03418729745

Q : 1 g. atom of nitrogen at STP represents

- ☐ A 6.02×10^{23} N_2 molecules
- ☐ B 22.4 dm^3 of N_2
- ☐ C 11.2 dm^3 of N_2
- ☐ D 28 g of N_2

SAEED MDCAT

SAEED MDCAT TEAM

SAEEDMDCAT

www.saeedmdcat.com

FB)SAEED MDCAT

03418729745

Q:

2.8 g of an unknown gas occupies 2.24 dm³ volume at standard temperature and pressure. The gas may be

- ☐ Carbon dioxide
- ☐ Carbon monoxide
- ☐ Oxygen
- ☐ Sulphur dioxide

SAEED MDCAT

SAEED MDCAT TEAM

SAEEDMDCAT

www.saeedmdcat.com

FB)SAEED MDCAT

03418729745

Q : The weight of 11.2 dm^3 of CO_2 at STP would be

☐ 88 g

☐ 44 g

☐ 32 g

☐ 22 g

SAEED MDCAT

SAEED MDCAT TEAM

SAEEDMDCAT

www.saeedmdcat.com

FB)SAEED MDCAT

03418729745

Q : The number of moles of CO_2 which contains 16g of oxygen

☐ 0.25

☐ 1.50

☐ 0.50

☐ 1.00

SAEED MDCAT

SAEED MDCAT TEAM

SAEEDMDCAT

www.saeedmdcat.com

FB)SAEED MDCAT

03418729745



22/10



25 min



100%

Q:

Efficiency of a chemical reaction is



Actual yield



Theoretical yield



Percentage yield



All of these

SAEED MDCAT

SAEED MDCAT TEAM



SAEEDMDCAT

www.saeedmdcat.com

FB)SAEED MDCAT

03418729745



23/10



25 min



100%

Q:

Which of the following are limitations of chemical equations



They do not tell about the conditions of reactions



Rate of reaction



Phase change involved



All of these

SAEED MDCAT

SAEED MDCAT TEAM



SAEEDMDCAT

www.saeedmdcat.com

FB)SAEED MDCAT

17

18

19

20

21

22

03418729745

Q :

A chemist is more interested about _____ to express the efficiency of a chemical process

- ☐ Theoretical yield
- ☐ Actual Yield
- ☐ %age yield
- ☐ Non-limiting reactant

SAEED MDCAT

SAEED MDCAT TEAM

SAEEDMDCAT

www.saeedmdcat.com

FB)SAEED MDCAT

03418729745

Q : Find the mass of iron which will be converted into its oxide by the action of 18g of steam as

☐ 42 g

☐ 56 g

☐ 4.2 g

☐ 28 g

SAEED MDCAT

SAEED MDCAT TEAM

SAEEDMDCAT

www.saeedmdcat.com

FB)SAEED MDCAT

24

26 27 28 29

30

03418729745

Q : Unit of concentration for representing the trace amount of substance is

- ☐ A Molarity
- ☐ B Mole fraction
- ☐ C Molality
- ☐ D Parts per million

SAEED MDCAT

SAEED MDCAT TEAM

SAEEDMDCAT

www.saeedmdcat.com

FB)SAEED MDCAT

03418729745

Q : When we dissolve 15.8g of KMnO_4 in 1000cm^3 of solution then the molarity of the solution is

- ☐ 0.1M
- ☒ 0.2M
- ☐ 0.01M
- ☐ 0.02M

SAEED MDCAT

SAEED MDCAT TEAM

SAEEDMDCAT

www.saeedmdcat.com

FB)SAEED MDCAT

24 25 26 28 29 30
03418729745

Q : The sum of mole fractions of components of a solution is always equal to

- ☐ A Zero
- ☐ B One
- ☐ C Two
- ☐ D 100

SAEED MDCAT

SAEED MDCAT TEAM

SAEEDMDCAT

www.saeedmdcat.com

FB)SAEED MDCAT

03418729745

23 : 51



29/30



25 min



Hint

Q : The molarity of an aqueous solution of NaOH containing 8 gm in 2 litre of solution is

A

0.1 M

B

0.2 M

C

0.25 M

D

0.15 M

SAEED MDCAT

SAEED MDCAT TEAM



SAEEDMDCAT

www.saeedmdcat.com

FB)SAEED MDCAT

24

25

26

27

28

29

30

03418729745

23 : 48



30/30



25 min



Hint

Q : A solution of 92g of ethanol, 96g methanol and 90g water has mole fraction of ethanol equal to



0.1



0.4



0.2



0.5

SAEED MDCAT

SAEED MDCAT TEAM



SAEEDMDCAT

www.saeedmdcat.com

FB)SAEED MDCAT

24

25

26

27

28

29

30

03418729745

SAEED MDCAT

TRUE HERO ALWAYS WINS

National MDCAT 2020 Results

190+ 35 Students

185+ 218 Students

180+ 677 Students



SAEED MDCAT
SAEED MDCAT TEAM
f SAEEDMDCAT

SAEED MDCAT
JOIN US FOR FREE
TO SECURE YOUR FUTURE

Designed by: Usman Arts

f SAEEDMDCAT

Unit-Wise - Test 5

Chemistry - Fundamental Concept

1-D	11-D	21-C
2-D	12-C	22-C
3-B	13-A	23-D
4-C	14-A	24-C
5-D	15-C	25-B
6-A	16-B	26-D
7-A	17-B	27-A
8-C	18-C	28-B
9-C	19-B	29-A
10-B	20-D	30-C

SAEED MDCAT

SAEED MDCAT TEAM



SAEEDMDCAT